

## Bis( $\mu$ -pyridazine-3-carboxylato- $\kappa^2$ O:O')-bis[aquadioxido(pyridazine-3-carboxylato- $\kappa^2$ N<sup>2</sup>,O)uranium(VI)] dihydrate

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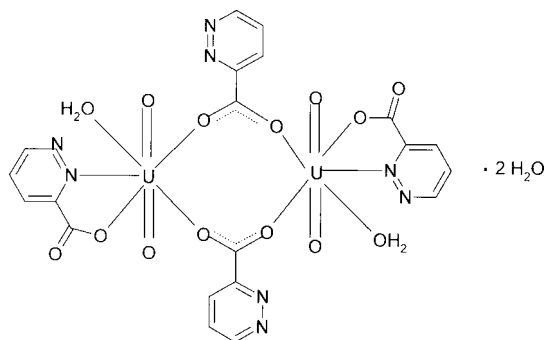
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 Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.013$  Å;  $R$  factor = 0.043;  $wR$  factor = 0.127; data-to-parameter ratio = 19.4.

The structure of the binuclear title complex,  $[\text{U}_2(\text{C}_5\text{H}_3\text{N}_2\text{O}_2)_4\text{O}_4(\text{H}_2\text{O})_2]\cdot 2\text{H}_2\text{O}$ , is composed of centrosymmetric dimers in which each  $\text{UO}_2^{2+}$  ion is coordinated by two ligand molecules. One donates its  $N,O$ -bonding group and the other donates both carboxylate O atoms. Each of the latter bridges adjacent uranyl ions. The coordination environment of the metal center is a distorted pentagonal bipyramid. The dimers are interconnected by  $\text{O}-\text{H}\cdots\text{O}$  hydrogen bonds between coordinated and uncoordinated water molecules and carboxylate O atoms. An intramolecular  $\text{O}-\text{H}\cdots\text{N}$  interaction is also present.

### Related literature

For the crystal structure of pyridazine-3-carboxylic acid hydrochloride, see: Gryz *et al.* (2003). For centrosymmetric dimeric molecules with a different bridging mode for the title ligand to calcium(II), see: Starosta & Leciejewicz (2007). For bond distances and angles in uranyl complexes with carboxylate ligands, see: Leciejewicz *et al.* (1995).



### Experimental

#### Crystal data

$[\text{U}_2(\text{C}_5\text{H}_3\text{N}_2\text{O}_2)_4\text{O}_4(\text{H}_2\text{O})_2]\cdot 2\text{H}_2\text{O}$   
 $M_r = 1104.50$   
 Monoclinic,  $C2/c$   
 $a = 25.660$  (5) Å  
 $b = 6.8330$  (14) Å  
 $c = 16.673$  (3) Å  
 $\beta = 96.73$  (3)°  
 $V = 2903.2$  (10) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 11.23$  mm<sup>-1</sup>  
 $T = 293$  (2) K  
 $0.19 \times 0.12 \times 0.07$  mm

#### Data collection

Kuma KM-4 four-circle diffractometer  
 Absorption correction: analytical (*CrysAlis RED*; Oxford Diffraction, 2008)  
 $T_{\min} = 0.234$ ,  $T_{\max} = 0.470$   
 4406 measured reflections  
 4266 independent reflections  
 2746 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.025$   
 3 standard reflections every 200 reflections  
 intensity decay: 1.0%

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.043$   
 $wR(F^2) = 0.127$   
 $S = 1.01$   
 4266 reflections  
 220 parameters  
 5 restraints  
 H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\text{max}} = 2.83$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -4.65$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O4}-\text{H42}\cdots\text{O22}^i$	0.81 (11)	2.23 (9)	2.933 (9)	146 (14)
$\text{O4}-\text{H41}\cdots\text{O22}$	0.83 (8)	1.98 (8)	2.803 (10)	175 (13)
$\text{O3}-\text{H31}\cdots\text{N11}$	0.82 (9)	1.94 (9)	2.754 (9)	170 (13)
$\text{O3}-\text{H32}\cdots\text{O4}^{\text{ii}}$	0.82 (6)	1.90 (9)	2.707 (12)	170 (13)

 Symmetry codes: (i)  $-x + \frac{1}{2}, -y + \frac{3}{2}, -z + 1$ ; (ii)  $x, -y + 1, z - \frac{1}{2}$ .

Data collection: *KM-4 Software* (Kuma, 1996); cell refinement: *KM-4 Software*; data reduction: *DATAPROC* (Kuma, 2001); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: RK2117).

### References

- Gryz, M., Starosta, W., Ptasiwicz-Bąk, H. & Leciejewicz, J. (2003). *J. Coord. Chem.* **56**, 1505–1511.  
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 Oxford Diffraction (2008). *CrysAlis RED*. Oxford Diffraction Ltd., Abingdon, England.  
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**supplementary materials**

*Acta Cryst.* (2009). E65, m94 [ doi:10.1107/S1600536808042219 ]

**Bis( $\mu$ -pyridazine-3-carboxylato- $\kappa^2 O:O'$ )bis[aquadioxido(pyridazine-3-carboxylato- $\kappa^2 N^2, O$ )uranium(VI)] dihydrate**

**J. Leciejewicz and W. Starosta**

**Comment**

The structure of the title compound is built of centrosymmetric dimeric molecules (Fig. 1) in which each  $UO_2^{2+}$  ion is coordinated by two ligand molecules: one chelates through its N,O-bonding group, the other donates only carboxylate O atoms which bridge two adjacent uranyl ions. The U ions, ligand molecules and coordinated water O atoms form a plane [r.m.s 0.1063 (2) Å]. The coordination around the uranyl ion is pentagonal bipyramidal. The equatorial plane of the pyramid [r.m.s 0.0674 (2) Å] consists of a water O3 atom, a hetro-cycle N21 atom, a carboxylate O21 atom donated by one ligand molecule, and two bridging O11 and O12 atoms, each donated by the other ligand molecules. Maximum shifts from the mean plane show the N21 [+0.1228 (2) Å] and the carboxylate O21 [-0.1056 (2) Å] atoms. The mean U—O bond distance in the  $UO_2^{2+}$  ion is 1.755 (8) Å], the O1—U—O2 angle is 178.2 (3)°. The U—O bond distances and bond angles within the equatorial plane fall in the range commonly observed in uranyl complexes with carboxylate ligands (Leciejewicz *et al.*, 1995). Bond lengths and bond angles within the ligand molecules agree well with those reported in the structure of the title ligand (Gryz *et al.*, 2003). The dimers interact by H bonds and form molecular sheet in which coordinated and solvent water molecules are the donors and carboxylate O atoms - the acceptors (Fig. 2). An intra-dimer H bond of 2.754 (9) Å is also observed.

**Experimental**

Hot aqueous solutions containing 2 mmol of pyridazine-3-carboxylic acid and 1 mmol of uranyl nitrate hexahydrate, respectively, were mixed and boiled for two hours with constant stirring and left to crystallize at room temperature. After few days, well formed green single crystals were found in the mother liquor. They were washed with cold ethanol and dried in air.

**Refinement**

C-bonded H atoms were placed in idealized positions and refined with a riding model approximation with C—H 0.93 Å and  $U_{iso}(H) = 1.2U_{eq}(C)$ . The positions of water H atoms were initially located from Fourier maps and refined with constraints on O—H distances and individual H isotropic displacement parameters.

A maximum peak of  $2.83 \text{ e}\text{\AA}^{-3}$  at  $1.06\text{\AA}$  and a deepest hole of  $-4.65 \text{ e}\text{\AA}^{-3}$  at  $0.96 \text{\AA}$  from the U1 atom were found on the final electron density map.

## Figures

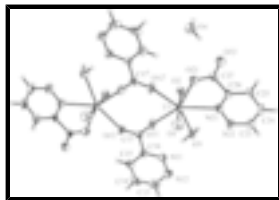


Fig. 1. View a dimer of title compound with the atom labelling scheme. Displacement ellipsoids are drawn at 50% probability level. H atoms are presented as a small spheres of arbitrary radius. Symmetry code: (i)-x+1, y, -z+1/2.



Fig. 2. Packing diagram of the crystal structure.

## Bis( $\mu$ -pyridazine-3-carboxylato- $\kappa^2$ O:O')bis[aquadioxido(pyridazine-3-carboxylato- $\kappa^2$ N<sup>2</sup>,O)uranium(VI)] dihydrate

### Crystal data

$[\text{U}_2(\text{C}_5\text{H}_3\text{N}_2\text{O}_2)_4\text{O}_4(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$

$M_r = 1104.50$

Monoclinic,  $C2/c$

Hall symbol:  $-C\ 2yc$

$a = 25.660\ (5)\ \text{\AA}$

$b = 6.8330\ (14)\ \text{\AA}$

$c = 16.673\ (3)\ \text{\AA}$

$\beta = 96.73\ (3)^\circ$

$V = 2903.2\ (10)\ \text{\AA}^3$

$Z = 4$

$F_{000} = 2032$

$D_x = 2.527\ \text{Mg m}^{-3}$

Mo  $K\alpha$  radiation

$\lambda = 0.71073\ \text{\AA}$

Cell parameters from 25 reflections

$\theta = 6\text{--}15^\circ$

$\mu = 11.23\ \text{mm}^{-1}$

$T = 293\ (2)\ \text{K}$

Block, light green

$0.19 \times 0.12 \times 0.07\ \text{mm}$

### Data collection

Kuma KM-4 four-circle diffractometer

Radiation source: Fine-focus sealed tube

Monochromator: Graphite

$T = 293\ (2)\ \text{K}$

$\omega/2\theta$  profile scans

Absorption correction: Analytical (CrysAlis RED; Oxford Diffraction, 2008)

$T_{\min} = 0.234$ ,  $T_{\max} = 0.470$

4406 measured reflections

4266 independent reflections

2746 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.025$

$\theta_{\max} = 30.1^\circ$

$\theta_{\min} = 1.6^\circ$

$h = -36 \rightarrow 35$

$k = -9 \rightarrow 0$

$l = 0 \rightarrow 23$

3 standard reflections

every 200 reflections

intensity decay: 1.0%

### Refinement

Refinement on  $F^2$

Least-squares matrix: Full

Secondary atom site location: Difmap

Hydrogen site location: Geom

$$R[F^2 > 2\sigma(F^2)] = 0.043$$

$$wR(F^2) = 0.127$$

$$S = 1.01$$

4266 reflections

220 parameters

5 restraints

Primary atom site location: Direct

H atoms treated by a mixture of independent and constrained refinement

$$w = 1/[\sigma^2(F_o^2) + (0.0893P)^2],$$

$$\text{where } P = (F_o^2 + 2F_c^2)/3$$

$$(\Delta/\sigma)_{\max} < 0.001$$

$$\Delta\rho_{\max} = 2.83 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -4.65 \text{ e } \text{\AA}^{-3}$$

Extinction correction: None

### Special details

**Geometry.** All s.u.'s (except the s.u. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell s.u.'s are taken into account individually in the estimation of s.u.'s in distances, angles and torsion angles; correlations between s.u.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell s.u.'s is used for estimating s.u.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
U1	0.388813 (9)	0.70410 (6)	0.214716 (15)	0.02772 (11)
O11	0.4687 (2)	0.7142 (12)	0.1613 (4)	0.0436 (16)
O21	0.3379 (2)	0.7086 (11)	0.3224 (3)	0.0425 (16)
C14	0.5457 (4)	0.7835 (16)	-0.0802 (6)	0.044 (2)
H14	0.5740	0.8037	-0.1092	0.052*
C16	0.5070 (3)	0.7471 (11)	0.0398 (5)	0.0267 (16)
O12	0.5529 (2)	0.7589 (12)	0.1677 (4)	0.0483 (19)
N21	0.2881 (3)	0.6906 (12)	0.1781 (4)	0.0365 (17)
N11	0.4595 (3)	0.7245 (12)	-0.0016 (4)	0.0345 (15)
C26	0.2583 (3)	0.6998 (13)	0.2373 (5)	0.0289 (15)
N12	0.4535 (3)	0.7342 (14)	-0.0813 (4)	0.0423 (19)
O1	0.3944 (2)	0.4495 (11)	0.2272 (4)	0.0420 (15)
C15	0.5519 (3)	0.7755 (15)	0.0024 (5)	0.0380 (19)
H15	0.5847	0.7884	0.0321	0.046*
C17	0.5096 (3)	0.7400 (12)	0.1300 (5)	0.0274 (16)
C13	0.4960 (5)	0.7605 (15)	-0.1178 (5)	0.046 (2)
H13	0.4916	0.7636	-0.1739	0.056*
O3	0.3712 (2)	0.6478 (13)	0.0725 (4)	0.0480 (19)
H32	0.361 (5)	0.537 (8)	0.060 (8)	0.072*
H31	0.398 (3)	0.655 (19)	0.050 (7)	0.072*
O22	0.2641 (2)	0.7074 (12)	0.3803 (4)	0.0478 (17)
O2	0.3827 (3)	0.9573 (12)	0.1990 (4)	0.0501 (17)
N22	0.2650 (3)	0.6893 (14)	0.0999 (4)	0.046 (2)

## supplementary materials

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C27	0.2882 (3)	0.7068 (14)	0.3202 (5)	0.0313 (15)
C25	0.2035 (3)	0.7050 (14)	0.2253 (6)	0.0371 (18)
H25	0.1833	0.7094	0.2681	0.045*
C23	0.2131 (4)	0.6890 (16)	0.0858 (6)	0.045 (2)
H23	0.1976	0.6786	0.0327	0.054*
C24	0.1812 (3)	0.7034 (18)	0.1457 (7)	0.051 (3)
H24	0.1450	0.7118	0.1331	0.061*
O4	0.3251 (3)	0.6999 (15)	0.5313 (5)	0.058 (2)
H42	0.311 (5)	0.75 (2)	0.567 (5)	0.086*
H41	0.306 (5)	0.71 (2)	0.488 (4)	0.086*

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
U1	0.01643 (13)	0.04731 (19)	0.02056 (14)	0.00034 (12)	0.00697 (8)	0.00059 (13)
O11	0.022 (3)	0.087 (5)	0.023 (3)	-0.010 (3)	0.007 (2)	-0.006 (3)
O21	0.027 (3)	0.082 (5)	0.020 (2)	-0.001 (3)	0.010 (2)	0.001 (3)
C14	0.048 (5)	0.054 (5)	0.033 (4)	-0.009 (4)	0.025 (4)	-0.002 (4)
C16	0.024 (3)	0.033 (5)	0.025 (3)	0.001 (2)	0.007 (3)	-0.003 (3)
O12	0.022 (3)	0.100 (6)	0.024 (3)	0.000 (3)	0.002 (2)	0.012 (3)
N21	0.021 (3)	0.065 (5)	0.025 (3)	0.001 (3)	0.008 (2)	0.003 (3)
N11	0.029 (3)	0.054 (5)	0.021 (3)	-0.002 (3)	0.005 (2)	0.002 (3)
C26	0.019 (3)	0.040 (4)	0.029 (3)	0.004 (3)	0.008 (2)	0.004 (3)
N12	0.045 (4)	0.060 (6)	0.021 (3)	-0.003 (4)	0.002 (3)	-0.003 (3)
O1	0.039 (3)	0.049 (4)	0.040 (3)	0.003 (3)	0.014 (3)	0.006 (3)
C15	0.026 (4)	0.060 (6)	0.030 (4)	-0.002 (4)	0.013 (3)	0.005 (4)
C17	0.023 (3)	0.038 (5)	0.023 (3)	0.005 (3)	0.009 (3)	0.005 (3)
C13	0.064 (6)	0.056 (7)	0.021 (4)	-0.007 (5)	0.010 (4)	0.000 (4)
O3	0.021 (3)	0.094 (6)	0.030 (3)	-0.003 (3)	0.009 (2)	0.000 (3)
O22	0.030 (3)	0.088 (5)	0.027 (3)	0.010 (3)	0.015 (2)	0.002 (3)
O2	0.057 (4)	0.047 (4)	0.048 (4)	-0.003 (3)	0.014 (3)	0.006 (3)
N22	0.027 (3)	0.088 (7)	0.024 (3)	0.001 (3)	0.003 (3)	0.005 (4)
C27	0.022 (3)	0.041 (4)	0.033 (4)	0.007 (3)	0.011 (3)	0.006 (4)
C25	0.021 (3)	0.046 (5)	0.046 (5)	-0.003 (3)	0.013 (3)	0.004 (4)
C23	0.032 (4)	0.069 (7)	0.034 (4)	0.000 (4)	-0.004 (3)	0.010 (4)
C24	0.014 (3)	0.074 (7)	0.064 (7)	0.000 (4)	-0.003 (3)	0.010 (6)
O4	0.041 (4)	0.100 (7)	0.033 (3)	0.012 (4)	0.010 (3)	0.008 (4)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

U1—O2	1.754 (8)	N21—N22	1.367 (10)
U1—O1	1.756 (8)	N11—N12	1.321 (10)
U1—O11	2.331 (6)	C26—C25	1.399 (10)
U1—O21	2.342 (5)	C26—C27	1.501 (11)
U1—O12 <sup>i</sup>	2.352 (6)	N12—C13	1.322 (13)
U1—O3	2.393 (6)	C15—H15	0.9300
U1—N21	2.588 (7)	C13—H13	0.9300
O11—C17	1.236 (9)	O3—H32	0.82 (6)

O21—C27	1.272 (9)	O3—H31	0.82 (9)
C14—C13	1.362 (15)	O22—C27	1.237 (9)
C14—C15	1.369 (12)	N22—C23	1.326 (11)
C14—H14	0.9300	C25—C24	1.382 (15)
C16—N11	1.338 (10)	C25—H25	0.9300
C16—C15	1.385 (10)	C23—C24	1.368 (15)
C16—C17	1.500 (11)	C23—H23	0.9300
O12—C17	1.218 (9)	C24—H24	0.9300
O12—U1 <sup>i</sup>	2.352 (6)	O4—H42	0.81 (11)
N21—C26	1.317 (9)	O4—H41	0.83 (8)
O2—U1—O1	178.2 (3)	N22—N21—U1	122.2 (5)
O2—U1—O11	88.9 (3)	N12—N11—C16	120.3 (7)
O1—U1—O11	90.7 (3)	N21—C26—C25	123.8 (8)
O2—U1—O21	93.1 (3)	N21—C26—C27	114.4 (6)
O1—U1—O21	88.1 (3)	C25—C26—C27	121.8 (7)
O11—U1—O21	152.6 (2)	N11—N12—C13	117.6 (8)
O2—U1—O12 <sup>i</sup>	90.3 (3)	C14—C15—C16	117.0 (8)
O1—U1—O12 <sup>i</sup>	91.4 (3)	C14—C15—H15	121.5
O11—U1—O12 <sup>i</sup>	79.1 (2)	C16—C15—H15	121.5
O21—U1—O12 <sup>i</sup>	73.5 (2)	O12—C17—O11	124.4 (7)
O2—U1—O3	90.4 (3)	O12—C17—C16	116.3 (7)
O1—U1—O3	87.8 (3)	O11—C17—C16	119.2 (7)
O11—U1—O3	72.4 (2)	N12—C13—C14	125.6 (8)
O21—U1—O3	134.8 (2)	N12—C13—H13	117.2
O12 <sup>i</sup> —U1—O3	151.5 (2)	C14—C13—H13	117.2
O2—U1—N21	86.0 (3)	U1—O3—H32	115 (9)
O1—U1—N21	93.3 (3)	U1—O3—H31	112 (9)
O11—U1—N21	144.2 (2)	H32—O3—H31	101 (10)
O21—U1—N21	63.2 (2)	C23—N22—N21	118.8 (7)
O12 <sup>i</sup> —U1—N21	136.3 (2)	O22—C27—O21	124.7 (8)
O3—U1—N21	72.1 (2)	O22—C27—C26	119.8 (7)
C17—O11—U1	173.0 (7)	O21—C27—C26	115.4 (6)
C27—O21—U1	128.7 (5)	C24—C25—C26	115.6 (8)
C13—C14—C15	116.8 (8)	C24—C25—H25	122.2
C13—C14—H14	121.6	C26—C25—H25	122.2
C15—C14—H14	121.6	N22—C23—C24	123.1 (9)
N11—C16—C15	122.6 (8)	N22—C23—H23	118.5
N11—C16—C17	116.3 (7)	C24—C23—H23	118.5
C15—C16—C17	121.0 (7)	C23—C24—C25	119.1 (8)
C17—O12—U1 <sup>i</sup>	150.0 (6)	C23—C24—H24	120.4
C26—N21—N22	119.4 (7)	C25—C24—H24	120.4
C26—N21—U1	118.2 (5)	H42—O4—H41	110 (12)

Symmetry codes: (i)  $-x+1, y, -z+1/2$ .

Hydrogen-bond geometry (Å, °)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
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## supplementary materials

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O4—H42…O22 <sup>ii</sup>	0.81 (11)	2.23 (9)	2.933 (9)	146 (14)
O4—H41…O22	0.83 (8)	1.98 (8)	2.803 (10)	175 (13)
O3—H31…N11	0.82 (9)	1.94 (9)	2.754 (9)	170 (13)
O3—H32…O4 <sup>iii</sup>	0.82 (6)	1.90 (9)	2.707 (12)	170 (13)

Symmetry codes: (ii)  $-x+1/2, -y+3/2, -z+1$ ; (iii)  $x, -y+1, z-1/2$ .



Fig. 1

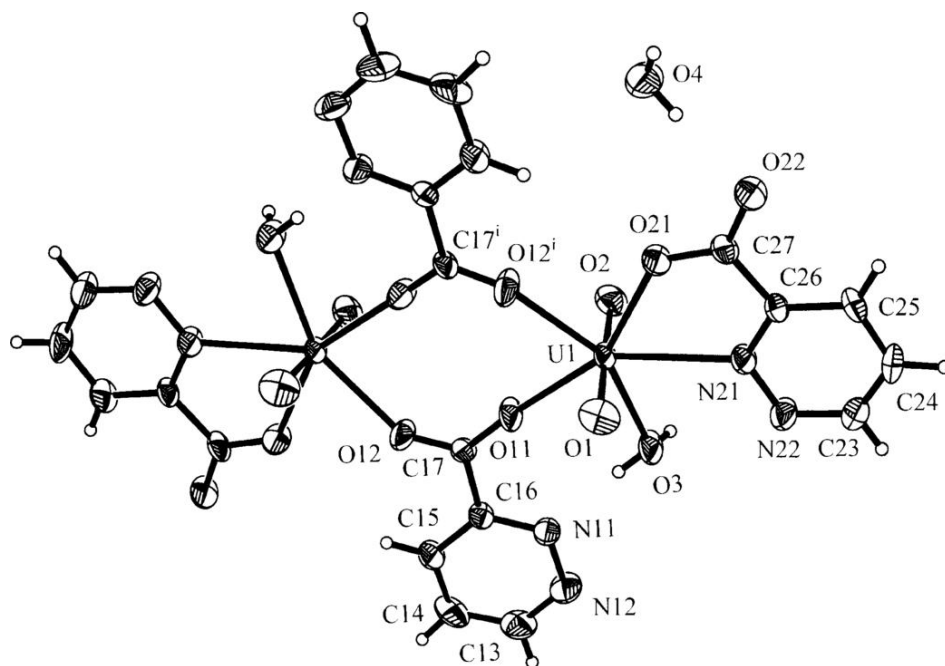


Fig. 2

